

Oxidation & Reduction | Notes

Definitions to learn

Reduction: Is the loss of electrons

Oxidation: Is the gain of electrons

Reducing agent: Electron donor

Oxidising agent: Electron receiver

<u>Metal</u>	<u>Heating with oxygen</u>	<u>Reaction with dilute HCl</u>
Potassium Sodium Calcium	Forms Oxide easily	Explosive
Magnesium Aluminium Zinc Iron	Slowly	Salt + Hydrogen gas
Hydrogen Copper Mercury	Slowly	No reaction
Silver Gold	NO reaction	No reaction

Sample Question

Identify what is oxidised and what is reduced in the following equations



