

Thermochemistry | Notes

Definitions

Heat of combustion: Heat change when one mole of a substance is burned in excess oxygen

Heat of formation: Heat change when one mole of a substance is made up from its elements in their standard state

Heat of solution: Heat change when one mole of a substance is dissolved in excess solution

Heat of reaction: Heat change that takes place according to a given equation.

Hess's law: Heat change in a chemical reaction depends only on the initial and final states and is independent of the path followed.

Endothermic reaction: Heat from a reaction taken in form its surroundings

Exothermic reaction: Heat from a reaction that's given out to its surroundings.

Sample Questions

Calculate the heat change of the following for the equation

